

## CHM 1045H – General Chemistry I

**Reference No:** 753385 – (T–R- 11:15 – 12:30 PM) – Room 1566

**Drop Date:** Last day to withdraw from a class with a W – Wednesday, 11/06/13

**Instructor:** Davia Hudson-Holness, Ph.D.

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**Office No:** 1547

**Web Page:** <http://faculty.mdc.edu/dholnes>

**Office Hours: Monday** 9:00AM – 10:00AM & 12:00PM – 2:00PM

**Tuesday** 7:15AM – 8:25AM & 9:50AM – 11:10AM

**Wednesday** 9:00AM – 10:00AM

**Thursday** 7:15AM – 8:25AM & 9:50AM – 11:10AM

**Friday** 9:00AM – 10:00AM

**Course Description:** CHM 1045 is the first term of a two-term sequence in inorganic chemistry. This course is primarily designed to prepare students for the higher level chemistry courses. This course will allow beginning chemistry students to explore the basis of important concepts in inorganic chemistry, such as atomic structure, chemical bonding, thermochemistry, stoichiometry, nomenclature, gases, etc., which will enable them to be successful in the second semester of General Chemistry, CHM 1046. The topics covered in this course meet MDC's general chemistry requirements, when taken in conjunction with CHM 1045L. This course is designed primarily for science, pre-health science and engineering students.

**Pre-requisites:** CHM 1025, or one year of High School Chemistry.

**Co-requisites:** CHM 1045L and MAC 1105 or higher

**Required Text:** CHEMISTRY – 3rd Edition by Julia Burdge, McGraw-Hill

**Recommended:** Student Solutions Manual by Burdge, Chang, Watkins and Hopkins  
3<sup>rd</sup> Edition, McGraw-Hill.

**Lecture Notes:** Course material, such as the syllabus, course schedule, and lecture slides will be available in ANGEL (<https://mycourses.mdc.edu/default.asp>). Please print lecture notes before class. All communication for the course will be via ANGEL. Please check it regularly since this is my way to contact you as a group.

**General Education Outcomes:** This class addresses the following [General Education Outcomes](#):

- Outcome No. 2 – Use quantitative analytical skills to evaluate and process numerical data.
- Outcome No. 3 – Solve problems using critical and creative thinking and scientific reasoning.
- Outcome No. 4 – Formulate strategies to locate, evaluate, and apply information.
- Outcome No. 8 – Use computer and emerging technologies effectively.
- Outcome No. 10 – Describe how natural systems function and recognize the impact of humans in the environment.

**Grading Policy:** Your grade in this course will be determined as follows:

**Tests and Exams** – There will be four tests of 100 points each. The lowest test grade will be dropped. There will be a final cumulative exam worth 200 points. ***There will be no makeup tests.***

**Report and Presentation** – Students will be assigned to groups and each group will be assigned a topic. A report and a 15 minute presentation will be required from each group at the end of the semester.

Graded Item	Points	Your Total Points	Grade	Grading Scale
Tests	300	$\geq 534$	A	89 – 100%
Report/ Presentation	100	$\geq 468$	B	78 – 88 %
		$\geq 402$	C	67 -77 %
Final Exam	200	$\geq 330$	D	55 – 66%
Total Points	600	$< 330$	F	54 – 0%

**Incompletes:** No Incompletes are given. A grade of “I” (Incomplete) will only be given under extraordinary circumstances: (a) that you missed the final exam for a good and verifiable reason; (b) that you have a passing average up to that point.

**Note:** A tentative, weekly schedule of topics and exam dates is given at the end of this handout.

**Attendance** – Attendance will be taken each class period in order to comply with government and college regulations. Student success in any class depends on their attendance to class and their professional behavior while in class, therefore, attendance is very important for the successful completion of this class. Professional conduct requires you to arrive on time to class and remain in class for the duration of the lecture session. Missed class notes and class materials are the responsibility of the student, and constitute no excuse for lack of performance in class or any of the tests or exams. **You may be purged from the class roster after 3 unexcused absences. If you have special circumstances, contact the instructor.**

**American Disabilities Act:** Any student, who because of disability may require special arrangements to meet the course requirements, should contact the instructor as soon as possible to make any necessary accommodations. Students requiring special accommodations in the classroom should present appropriate verification of their disability from the ACCESS Department (Room 1180, Phone Number 305-237-3072)

**Classroom Decorum:** Smoking, eating or drinking are not allowed in MDC classrooms. All students are expected to behave professionally and courteously while in class. You will receive one warning for any disruptions during class (e.g. unruly behavior, etc.). After this warning you will be asked to leave the classroom for the day and you will be counted as absent from the lecture. Additionally, students will be asked to leave the classroom if any of the following rules are violated during class:

- All cell phones must be **turned off or placed in silent mode and put away** prior to entering the classroom
- Use of cell phones, including earpieces, are strictly prohibited during class
- Cell phones cannot be used as calculators during a test, and text messaging is strictly forbidden while in class, and especially during a test or exam.

**Academic Dishonesty Policy:** If you are suspected of cheating, plagiarism, or any other type of academic dishonesty as outlined in MDC procedure 4074 of the [Students Rights and Responsibilities Handbook](#), you will be subjected to procedural due process as outlined in procedure 4074 of the same handbook. Honors students should refer to the Honors College Code of Conduct.

**Additional Help:** Additional help may be obtained in The Natural Science Tutoring Center (Room 2221) or The STEM Center (Room 2214).

**Electronic Calculators:** Electronic calculators are essential for the successful completion of homework assignments, as well as for tests and exams, and must be brought to class every time the class meets. You are required to bring a single line, non-programmable calculator (preferably solar powered) every time you attend this class (especially important during tests and exams).

*Programmable and graphing calculators are not allowed in this class. No two students are allowed to share a calculator, and you cannot use your cell phone as a calculator during a test or an exam. Students who fail to comply with these regulations will be asked to leave the classroom and will earn a grade of zero (0) in that test or exam.*

**Disclaimer:** The instructor reserves the right to change the order of delivery of topics and the date of test and exams, to accommodate the needs of the students in the class.

**Emergencies:** In case of an emergency, such as hurricanes or inclement weather, contact the instructor at 305-237-3113. A recording with instructions will be available at this phone number. If a test is scheduled for the day of the emergency, the test will be given the next time the class meets. Miami Dade College has an emergency number that you can use to inquire about the status of the College during an emergency on a 24 hour basis. The number is 305- 237-7500.

## CHM 1045H - Tentative Weekly Schedule

### Fall Semester 2013

Week	Dates	Topics	Chapter No.
1	08/26 – 08/30	First Day Handouts Chemistry: The central Science	<b>1</b>
2	09/02 – 09/06	Chemistry: The central Science Atoms, Molecules and Ions	<b>1</b> <b>2</b>
3	09/09 – 09/13	Atoms, Molecules and Ions (continued) <b>TEST I</b>	<b>2</b>
4	09/16 – 09/20	Stoichiometry	<b>3</b>
5	09/23 – 09/27	Stoichiometry Reactions in aqueous solutions	<b>3</b> <b>4</b>
6	09/30 – 10/04	Reactions in aqueous solutions (continued) <b>TEST II</b>	<b>4</b>
7	10/07 – 10/11	Thermodynamics	<b>5</b>
8	10/14 – 10/18	Thermodynamics (continued) Quantum Theory and the Electronic Structure of the Atom	<b>5</b> <b>6</b>
9	10/21 – 10/25	Quantum Theory and the Electronic Structure of the Atom (continued) <b>TEST III</b>	<b>6</b>
10	10/28 – 11/01	Electronic Configuration and the Periodic Table	<b>7</b>
11	11/04 – 11/08	Electronic Configuration and the Periodic Table (continued) Chemical Bonding I	<b>7</b> <b>8</b>
12	11/11 – 11/15	Chemical Bonding I (continued) <b>TEST IV</b>	<b>8</b>
13	11/18 – 11/22	Chemical Bonding II	<b>9</b>
14	11/25 – 11/29	Chemical Bonding II (continued) <b>Holiday</b>	<b>9</b>
15	12/02 – 12/06	Gases	<b>11</b>
16	12/09 – 12/13	Presentations Review for Final Exam	
17	<b>12/19</b>	<b>FINAL EXAM (Cumulative)</b> <b>11:15 AM – 12:45 PM</b> <b>Room 1566</b>	